Using Moles to Balance Equations

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|  | **Reactants** | **Products** |
| **1** | Sb = 488g  Cl2 =426g | SbCl3=914g |
| **2** | Mg=24g  O2=16g | MgO=40g |
| **3** | CaCl2=333g | Ca =120g  Cl2=213g |
| **4** | NaClO3 =106.5g | NaCl=58.5g  O2=48g |
| **5** | Fe=28g  HCl =36.5g | FeCl2=63.5g  H2=1g |
| **6** | CuO =19.875g  H2 =0.5g | Cu=15.875g  H2O=4.5g |
| **7** | Al =27g  H2SO4=147g | Al2(SO4)3=171g  H2=3g |

Task: Use the information below to construct a balanced symbol equation for each reaction. There are tables for you to lay your work out in.

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| **Equation from Q** |  |
| **Masses** |  |
| **Mr** |  |
| **Number of Moles** |  |
| **Ratio** |  |
| **Simplified Ratio** |  |

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| --- | --- |
| **Equation from Q** |  |
| **Masses** |  |
| **Mr** |  |
| **Number of Moles** |  |
| **Ratio** |  |
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